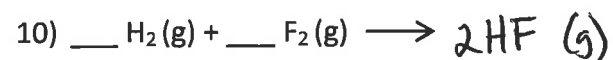
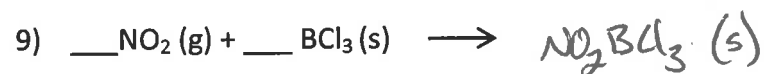
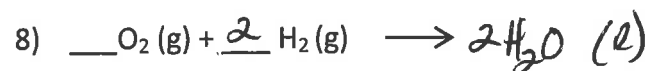
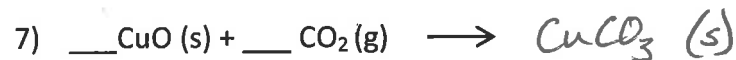
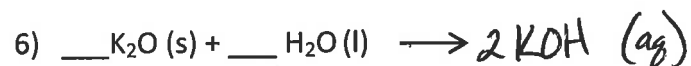
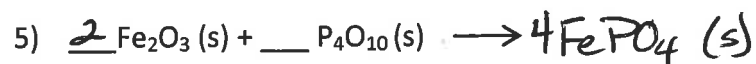
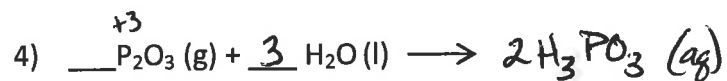
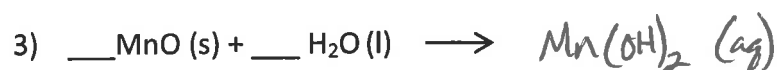
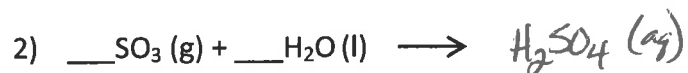
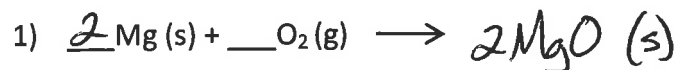


Key ✓

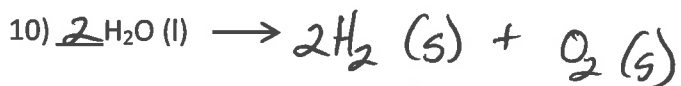
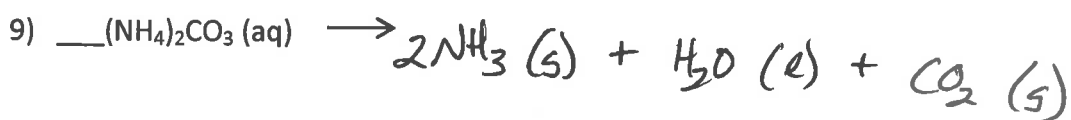
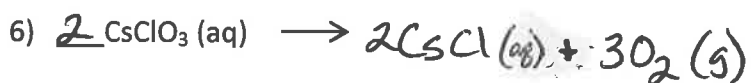
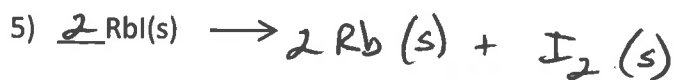
## AP Chemistry

### Types of Reactions Practice Packet

#### Synthesis



### Decomposition



Single Displacement (could be no rxn)

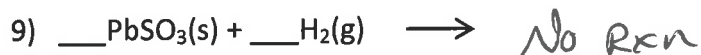
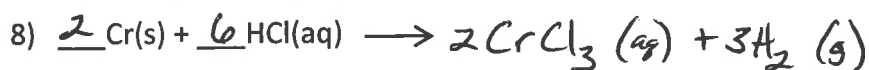
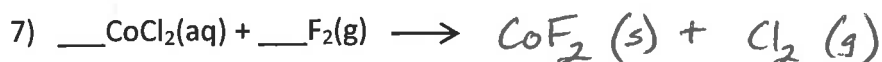
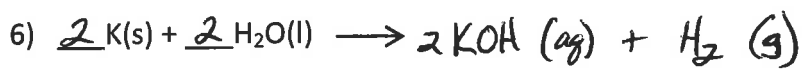
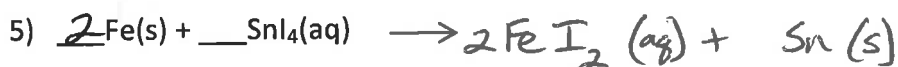
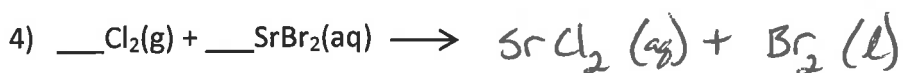
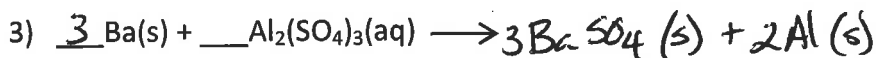
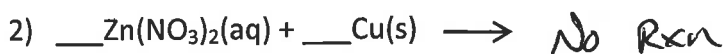
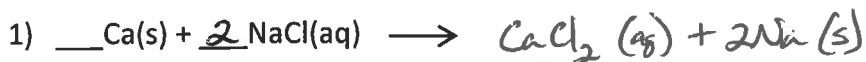
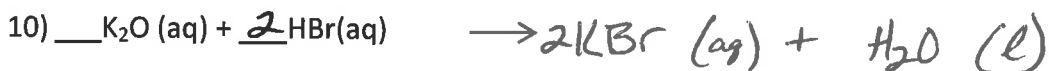
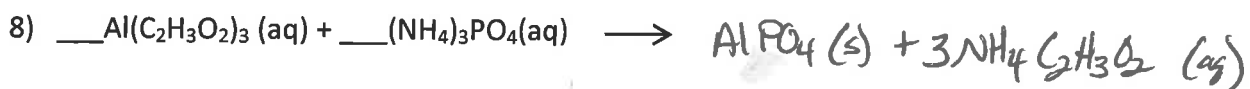
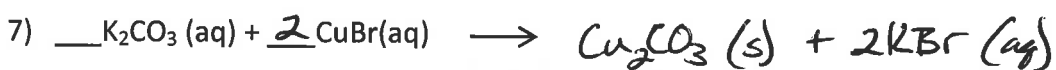
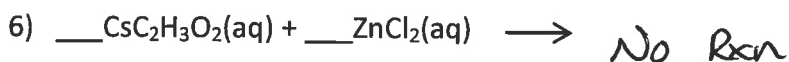
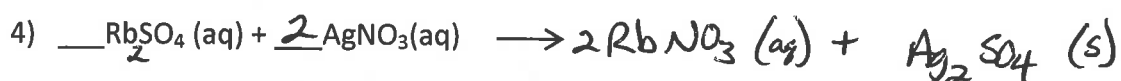
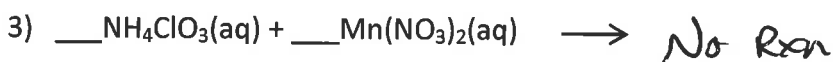
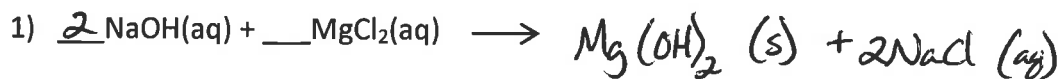


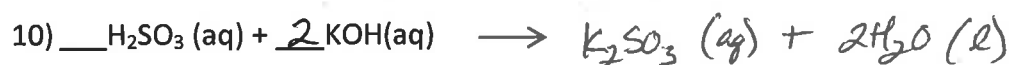
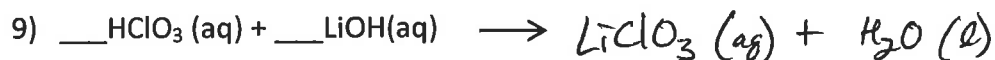
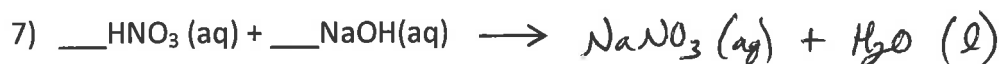
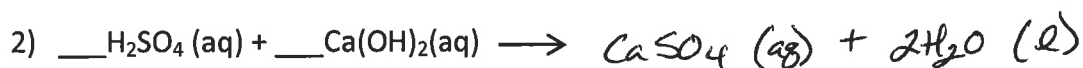
Table J  
Activity Series\*\*

Most	Metals	Nonmetals	Most
↓	Li	F <sub>2</sub>	↓
	Rb	Cl <sub>2</sub>	
	K	Br <sub>2</sub>	
	Cs	I <sub>2</sub>	
	Ba		
	Sr		
	Ca		
	Na		
	Mg		
	Al		
	Ti		
	Mn		
	Zn		
	Cr		
	Fe		
	Co		
	Ni		
	Sn		
	Pb		
**H <sub>2</sub>			
Cu			
Ag			
Au			
Least			Least

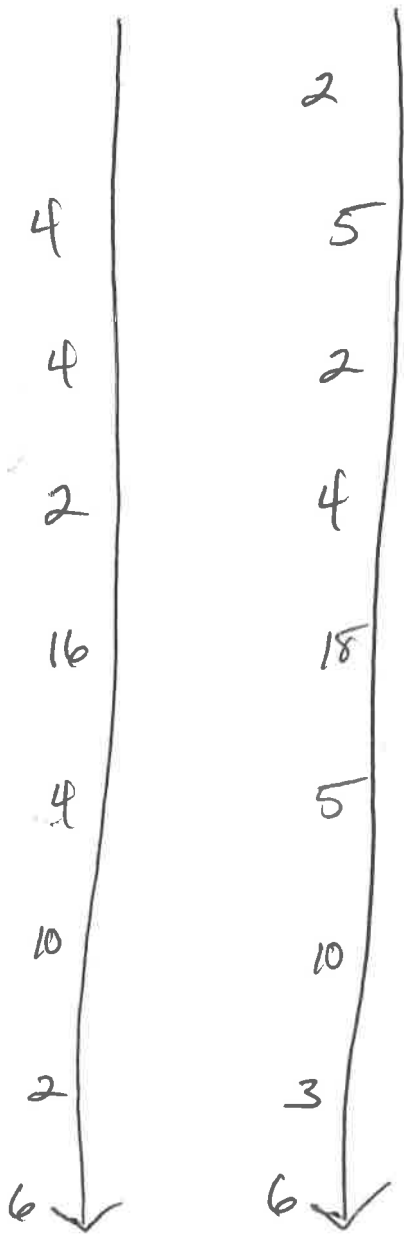
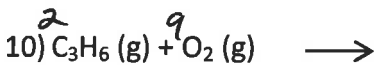
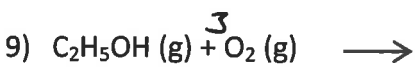
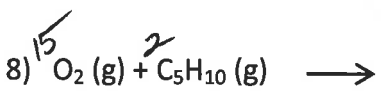
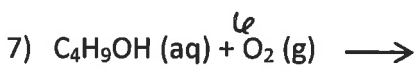
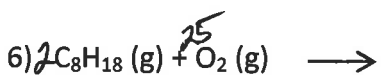
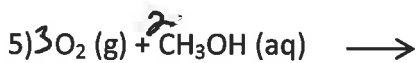
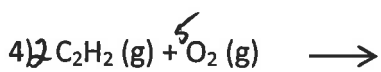
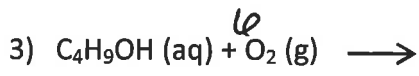
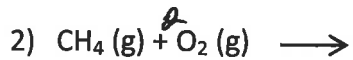
Double Displacement (could be no rxn)



### Acid-Base

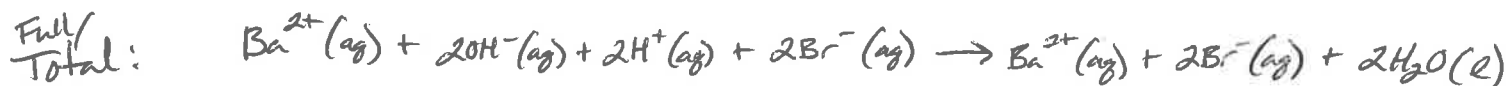
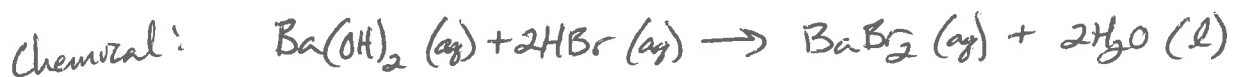


Combustion



### Net-Ionic Equations

- 1) Write the chemical, full ionic, and net ionic equations for a reaction between barium hydroxide and hydrobromic acid.

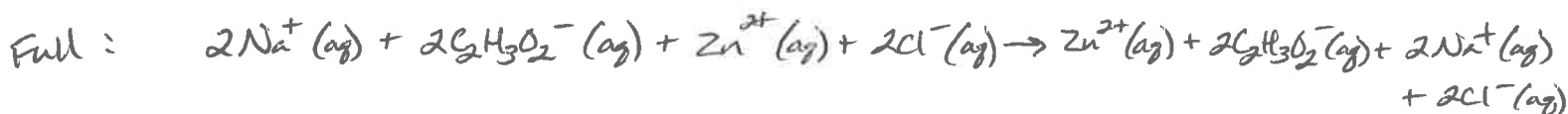


Net:

or



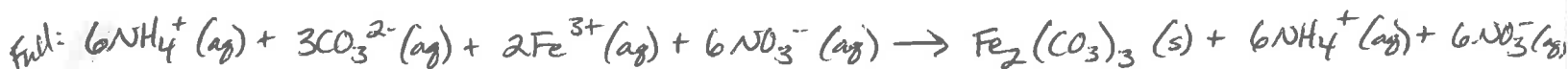
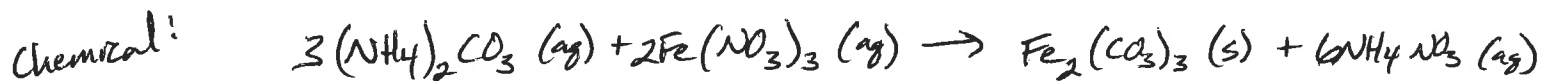
- 2) Write the chemical, full ionic, and net ionic equations for a reaction between sodium acetate and zinc chloride.



Net:

No reaction

- 3) Write the chemical, full ionic, and net ionic equations for a reaction between ammonium carbonate and iron (III) nitrate.



Net:

